

This syllabus is provided as a general informational guide. Some of the information may vary depending on the specific course section and instructor. Different sections of the same course may require different textbooks. Verify the section specific textbook information in the CUNY's Academic Course Schedule Web Page. Modifications of the grading system presented here will be communicated by the instructors of the sections when they meet the class.

BOROUGH OF MANHATTAN COMMUNITY COLLEGE

City University of New York

Department of Science

Title of Course: Quantitative Analysis

CHE 205

Semester

Credits 4

Class hours 3

Lab hours 6

Instructor

Name:

Office:

Tel:

Email:

A. COURSE DESCRIPTION:

This course discusses the principles of classical and instrumental techniques in analytical chemistry. Laboratory experiments include gravimetric, volumetric and instrumental methods of analysis.

Required in A.S. Forensic Science

Elective in A.S. Science

Prerequisites/Co-requisites

CHE 202, MAT 206 or Permission of the Department of Science

Basic Skills Prerequisites: ACR 095, ENG 095 or ESL 095, and MAT 056

Student Learning Outcomes

Students will

- a. understand the chemical principles important to analytical chemistry.
- b. be able to determine the accuracy and precision of experimental data.
- c. have laboratory skills necessary to solve analytical problems quantitatively.
- d. be able to obtain high quality analytical data.

Required Text & Readings

Skoog, D.A., West, D.M., Holler, F.J., and Crouch, S.R. *Fundamentals of Analytical Chemistry 9th Edition* Thomson/Brooks/Cole, 2004

Other Resources

Use of Technology (if applicable)

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Evaluation & Requirements of Students

There will be a minimum of four one hour examinations, a comprehensive final examination and fulfillment of all laboratory assignments.

College Attendance Policy

At BMCC, the maximum number of absences is limited to one more hour than the number of hours a class meets in one week. For example, you may be enrolled in a three-hour class. In that class, you would be allowed 4 hours of absence (not 4 days). In the case of excessive absences, the instructor has the option to lower the grade or assign an F or WU grade.

Academic Adjustments for Students with Disabilities

Students with disabilities who require reasonable accommodations or academic adjustments for this course must contact the Office of Services for Students with Disabilities. BMCC is committed to providing equal access to all programs and curricula to all students.

BMCC Policy on Plagiarism and Academic Integrity Statement

Plagiarism is the presentation of someone else's ideas, words or artistic, scientific, or technical work as one's own creation. Using the idea or work of another is permissible only when the original author is identified. Paraphrasing and summarizing, as well as direct quotations, require citations to the original source. Plagiarism may be intentional or unintentional. Lack of dishonest intent does not necessarily absolve a student of responsibility for plagiarism.

Students who are unsure how and when to provide documentation are advised to consult with their instructors. The library has guides designed to help students to appropriately identify a cited work. The full policy can be found on BMCC's web site, www.bmcc.cuny.edu. For further information on integrity and behavior, please consult the college bulletin (also available online).

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Outline of Topics

Lecture

Week 1	Topic	Page
Chapter 1	Nature of Analytical Chemistry	
1A	The Role of Analytical Chemistry	3 – 4
1B	Quantitative Analytical Methods	4 – 5
1C	Typical Quantitative Analysis	5 – 10
1D	Integral Role for Chemical Analyses	10 - 11
Chapter 2	Chemicals, Apparatus and Unit Operations of Analytical Chemistry	
2D-4	Sources of Error in Weighing	27 – 30
2G-2	Effect of Temperature on Volume Measurements	39 – 40
Week 2		
Chapter 5	Errors in Chemical Analyses	
5A-1	Mean and Median	92 – 92
5a-2	Precision	92 - 93
5A -3	Accuracy	93 – 94
5A-4	Types of Error in Experimental Data	94 – 95
5B	Systematic Errors	95 – 99
Chapter 6	Random Errors in Chemical Analysis	
6A	The Nature of Random Errors	105 - 107
Week 3		
Chapter 4	Calculations Used in Analytical Chemistry	
4A-2	Distinction Between Mass and Weight	72 - 73
4A-5	Calculations in Moles and Millimoles	74 - 76
4B-1	Concentrations of Solutions	76 – 83
Chapter 12	Precipitation Gravimetry	
12A-1	Properties of Precipitates	315 – 315
12A-2	Particle Size and Filterability of Precipitates	316 – 317
12A-3	Colloidal Precipitates	317 – 320
12A-4	Crystalline Precipitates	320 – 321
12A-5	Coprecipitation	321 – 324
12A-6	Precipitation from Homogenous Solutions	324 – 324
12A-7	Drying and Igniting of Precipitates	324 - 326

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Chapter 12	Precipitation Gravimetry	
12B	Calculation of Results from Gravimetric Data	326 – 329
12C	Applications of Gravimetric Methods	329 – 333
Chapter 13	Precipitation Titrimetry	
13A-1	Equivalence Points and End Points	338 – 340
13A-2	Primary Standards	340 – 340
13B	Standard Solutions	340 – 341
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Chapter 13	Precipitation Titrimetry	
13C	Volumetric Calculations	341 – 349
Chapter 9	Aqueous Solutions and Chemical Equilibria	
9A-1	Solutions of Electrolytes	228 – 229
9A-2	Acids and Bases	229 – 231
9B-6	Applying Acid Base Dissociation Constants	242 – 251
9C	Buffer Solutions	251 – 260
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17A-1	Complexation Equilibria	450 – 452
17A-2	Formation of Insoluble Species	453 – 453
17A-3	Ligands that Can Protonate	453 - 454
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Chapter 10	Effect of Electrolytes on Chemical Equilibria	
10A	Effect of Electrolytes on Chemical Equilibria	267 -269
10A-1	Effect of Ionic Charges on Equilibria	269 -269
10A-2	Effect of Ionic Strength	269 – 270
10A-3	The Salt Effect	271 – 271
10B	Activity Coefficients	271 – 272
10B-1	Properties of Activity Coefficients	272 – 273
10B-2	The Debye-Huckel Equation	273 – 275
10B-3	Equilibrium Calculations Using Activity Coefficients	275 – 278
10B-4	Omitting Activity Coefficients in Equilibrium Calculations	278- 278
Chapter 11	Solving Equilibrium Problems for Complex Systems	
11A	Solving Multiple Equilibrium Problems By a Systematic Method	282 – 282
11A-1	Mass Balance Equations	282 - 283
11a-2	Charge Balance Equations	284 – 285

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11A-4	Using Approximations to Solve Equilibrium Calculations	286 - 287
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Chapter 11	Solving Equilibrium Problems for Complex Systems	
11B	Calculating Solubilities by the Systematic Method	287 – 287
11B-1	The Solubility of Metal Hydroxides	288 – 291
11B-2	The Effect of pH on Solubility	291 – 294
11B-3	The Effect of Undissociated Solutes on Precipitation Calculations	294 – 299
11C	Separation of Ions By Control of the Concentration of the Precipitating Agent	300 – 300
11C-1	Calculation of the Feasibility of Separations	300 – 301
11C-2	Sulfide Separations	301 - 304
Chapter 13	Precipitation Titrimetry	
13D	Gravimetric Tirtimetry	349 – 350
13D-1	Calculations Associated with Weight Tritrations	350 – 350
13D-2	Advantages of Weight Titrations	350 – 350
13E	Titration Curves in Titrimetric Methods	350 – 351
13E-1	Types of Titration Curves	351 – 351
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13F-1	Precipitation Titration Curves Involving Silver Ion	353 – 356
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13F-3	Indicators for Argentometric Titrations	358 – 362
13F-4	Applications of Standard Silver Nitrate Solutions	362 363
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Chapter 14	Principles of Neutralization Titrations	
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14B	Titration of Strong Acids and Strong Bases	372 – 372
14B-1	Titrating a Strong Acid with a Strong Base	372 – 376
14B-2	Titrating a Storng Base with a Strong Acid	376 – 378
14C	Titration Curves for Weak Acids	378 – 381
14C-1	The Effect of Concentration	382 – 382
14C-2	The Effect of Reaction Completeness	382 – 382
14C-3	Choosing an Indicator: The Feasibility of Titration	382 – 383
14D	Titration Curves for Weak Bases	383 – 387
14E	The Composition of Solutions During Acid Base Titrations	387 – 390
Week 10	Titration Curves for Complex Acid Base Systems	

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15A	Mixtures of Strong and Weak Acids of Strong and Weak Baes	395 – 399
15B	Polyfunctional Acids and Bases	399 – 399
15B-1	The Phosphoric Acid System	399 – 400
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15D	Calculation of pH Solutions of NaHA	403 – 407
15E	Titration Curves for Polyfunctional Acids	407 – 415
15F	Titration Curves for Polyfunctional Bases	416 – 417
15G	Titration Curves for Amphiprotic Species	417 – 419
15H	The Compositions of Solutions of a Polyprotic Acid as a Function of pH	419 – 421
Chapter 16	Applications of Neutralization Titrations	
16A	Reagents for Neutralization Titrations	429 – 429
16A-1	Preparation of Standard Acid Solutions	429 – 429
16A-2	The Standardization of Acids	429 – 432
16A-3	Preparation of Standard Solutions of Bases	432 – 434
16A-4	The Standardization of Bases	434 – 434
16B	Typical Applications fo Neutralization Titrations	435 – 435
16B-1	Elemental Analysis	435 – 438
16B-2	The Determination of Inorganic Substances	438 - 441
16B-3	The Determination of Organic Functional Groups	441 – 443
16B-4	The Determination of Salts	
Week 11		
Chapter 17	Complexation Reactions and Titrations	
17B	Titrations with Inorganic Complexing Agents	455 – 457
17C	Organic Complexing Agents	457 – 458
17D	Amino Carboxylic Acid Titrations	458 – 458
17D-1	Ethylenediaminetetraacetic Acid	458 – 461
17D-2	Complexes of EDTA and Metal Ions	461 – 461
17D-3	Equilibrium Calculations Involving EDTA	461 – 466
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Chapter 17	Complexation Reactions and Titrations	
17D-4	EDTA Titration Curves	466 – 472
17D-5	The Effect of Other Complexing Agents on EDTA Titration Curves	472 – 475
17D-6	Indicators for EDTA Titrations	475 – 478
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17D-8	The Scope of EDTA Titrations	479 – 481
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18A	Characterizing Oxidation-Reduction Reactions	490 – 491
18A-1	Comparing Redox Reactions to Acid Base Reactions	491 – 493
18A-2	Oxidation-Reduction Reactions in Electrochemical Cells	493 – 494
18B	Electrochemical Cells	494 – 495
18B-1	Cathodes and Anodes	496 – 496
18B-2	Types of Electrochemical Cells	496 – 498
18B-3	Representing Cells Schematically	498 – 499
18B-4	Currents in Electrochemical Cells	499 – 499
18C	Electrode Potentials	499 – 501
18C-1	Sign Convention for Cell Potentials	502 – 504
18C-2	The Standard Hydrogen Reference Electrode	504 – 505
18C-3	Electrode Potential and Standard Electrode Potential	505 – 508
18C-4	Additional Implications of the IUPAC Sign Conventions	508 - 508
18C-5	Effect of Concentration on Electrode Potentials: The Nernst Equation	508 – 510
18C-6	The Standard Electrode Potential E°	511 - 516
18C-7	Limitations to the Use of Standard Electrode Potentials	516 – 518
Week 13		
Chapter 19	Applications of Standard Electrode Potentials	
19A	Calculating Potentials of Electrochemical Cells	522 – 530
19B	Determining Standard Potentials Experimentally	530 – 532
19C	Calculating Redox Equilibrium Constants	532 – 538
19D	Constructing Redox Titration Curves	538 – 539
19D-1	Electrode Potentials During Redox Titrations	539 – 541
19D-2	The Titration Curve	542 – 550
19D-3	Effect of Variables on Redox Titration Curves	551 – 552
19E	Oxidation Reduction Indicators	552 – 552
19E-1	General Redox Indicators	552 – 555
19E-2	Specific Indicators	555 – 555
19F	Potentiometric End Points	555 - 555
Chapter 20	Applications of Oxidation-Reduction Titrations	
20A	Auxilliary Oxidizing and Reducing Agents	560 – 561
20A-1	Auxilliary Reducing Reagents	561 – 562
20A-2	Auxilliary Oxidizing Reagents	562 – 562
20B	Applying Standard Reducing Agents	562 – 566
20C	Applying Standard Oxidizing Agents	566 – 582
Week 14		
Chapter 24	Introduction to Spectrochemical Methods	
24A	Properties of Electromagnetic Radiation	711 – 714
24B	Interaction of Radiation with Matter	714 – 717
24C	Radiation Absorption	718 – 734

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24D	Emission of Electromagnetic Radiation	734 – 739
Chapter 26	Molecular Absorption Spectroscopy	
26a	Ultraviolet and Visible Molecular Absorption Spectroscopy	784 - 807
26C	Infrared Absorption Spectroscopy	811 - 819
Chapter 27	Molecular Fluorescence Spectroscopy	
27A	Theory of Molecular Fluorescence	825 – 828
27B	Effect of Concentration on Fluorescence Intensity	829 – 830
27C	Fluorescence Instruments	830 – 831
27D	Applications of Fluorescence Methods	831 - 834
Chapter 28	Atomic Spectroscopy	
28A	Origins of Atomic Spectra	840 – 843
28B	Production of Atoms and Ions	843 – 854
28C	Atomic Emission Spectroscopy	856 – 858
28D	Atomic Absorption Spectroscopy	858 – 868
28E	Atomic Fluorescence Spectroscopy	868 – 868
28G	Atomic Mass Spectrometry	868 - 871

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Experiment	Title
1	Laboratory Safety
2	Gravimetric Determination of Sulfate as Barium Sulfate
3	Gravimetric Determination of Chloride
4	Gravimetric Determination of Calcium
5	Titration of Chloride with Silver Nitrate
6	Preparation and Standardization of Acids and Bases
7	Acid Base Titration Analysis
8	Determination of Magnesium by Complex Formation Titration
9	Preparation and Standardization of Potassium Permanganate
10	Determining the Percentage of Iron in Iron Ore
11	Iodimetric Titration of Vitamin C
12	Oxidation-Reduction Titration
13	Spectrophotometric Determination of Iron in Vitamins
14	Spectrophotometric Determination of an Equilibrium Constant
15	Examination Week